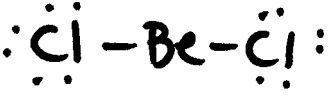
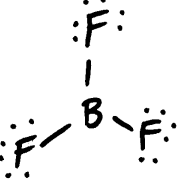
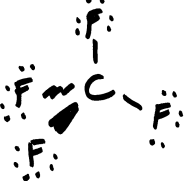
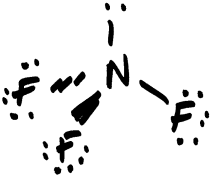
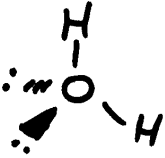
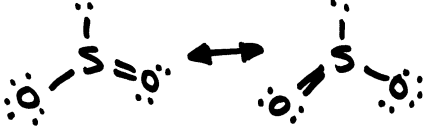
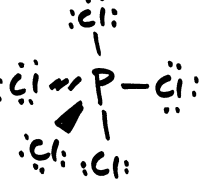
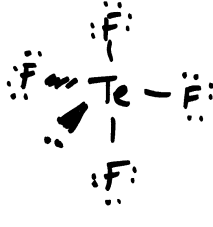
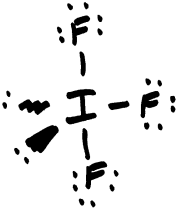
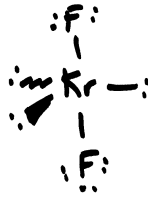
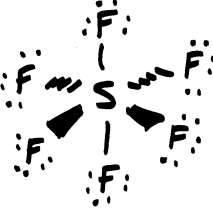
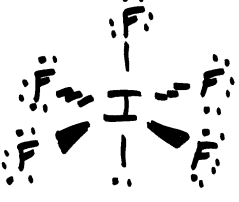
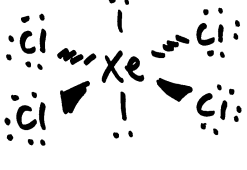


## Molecular Structure Flash Cards

KCl	K - Cl	linear 180° polar no hybridization (s)
BeCl <sub>2</sub>		linear 180° nonpolar sp
BF <sub>3</sub>		trigonal planar 120° nonpolar sp <sup>2</sup>
CF <sub>4</sub>		tetrahedral 109.5° nonpolar sp <sup>3</sup>
NF <sub>3</sub>		trigonal pyramidal 107° polar sp <sup>3</sup>
H <sub>2</sub> O		bent 104.5° polar sp <sup>3</sup>
SO <sub>2</sub>		bent <120° polar sp <sup>2</sup>
PCl <sub>5</sub>		trigonal bipyramidal 90°, 120° nonpolar dsp <sup>3</sup>

TeF <sub>4</sub>		seesaw, irregular tetrahedron, or bi-sphenoid <90°, <120° polar dsp <sup>3</sup>
IF <sub>3</sub>		T-shaped (planar) <90° polar dsp <sup>3</sup>
KrF <sub>2</sub>		linear 180° nonpolar dsp <sup>3</sup>
SF <sub>6</sub>		octahedral 90° nonpolar d <sup>2</sup> sp <sup>3</sup>
IF <sub>5</sub>		square pyramidal <90° polar d <sup>2</sup> sp <sup>3</sup>
XeCl <sub>4</sub>		square planar 90° nonpolar d <sup>2</sup> sp <sup>3</sup>