

Unit 3 Normality

** Disproportionation means an element is both oxidized and reduced.

** Normality is the number of equivalents (H^+ , OH^- for acid/base reactions or what supplies one mol of electrons in a redox) of a substance dissolved in one liter of solution.

Normality: $1 \text{ N} = 1 \text{ mol H}^+ / 1 \text{ liter}$ {ex. **3 M H_2SO_4** = 3 mol H_2SO_4 / 1 liter, but it has two hydrogens so normality = [(3 mol H_2SO_4) (**2 mol H^+**)] / 1 liter = **6 N H^+ in H_2SO_4** }

Molarity and Normality are only equal, if there is just one H^+ or one OH^- in the compound (or one mol of electrons in a redox).

End of Notes